

PHASES OF MATTER TEST REVIEW ①

Name:

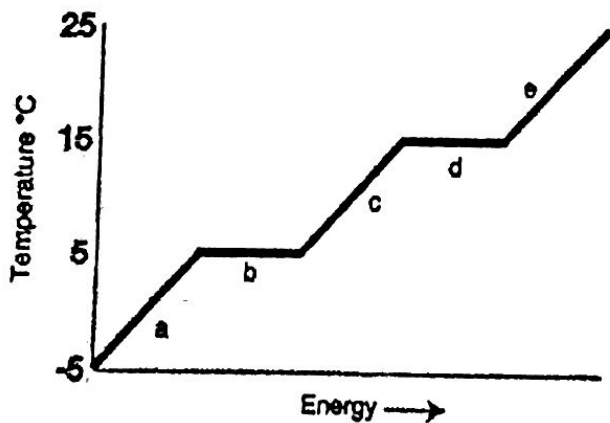
Define:

Endothermic:

Exothermic:

Complete the following table and label the phase diagram below.

Phase Change	From	To	Exothermic or Endothermic	Temperature
Melting	Solid	Liquid		
Freezing				
Boiling				
Condensing				
Vaporizing				
Evaporating				
Deposition				skip
Sublimation				skip



- 1.. What happens to kinetic energy (movement of the atoms) when temperature is increased ?
2. In what state of matter do molecules have the most energy? _____
3. In what state of matter do molecules have the least energy? _____
4. At what point on the graph does the substance have the most kinetic energy? _____

5. . At what point does the substance have the least kinetic energy? _____
- 6.. What is the melting point of this substance? _____
7. What other phase change shares the same temperature? _____
8. What state of matter is represented by a? _____ b? _____ c? _____
- 9.. What is happening to the temperature at point B? _____
10. What letter represents the vaporization of this substance? _____
11. What is the vaporization point of this substance? _____
12. What letter represents the warming of a solid? _____
13. Why does the temperature plateau during a phase change? _____

Refer to the following chart showing the melting and boiling point of different substances when answering the questions below.

Substance	Melting Point (Celsius)	Boiling Point (Celsius)
Water	0°	100°
Nitrogen	-210°	-195.8°
Mercury	-38.4°	357°
Iron	1536°	3000°

14. Which of the substances are liquids at room temperature (20° C)?

15. Which of the substances are gases at 115° C ?

16. Which of the following substances are solid at room temperature?

Temperature Review:

1. What is absolute zero? _____

Convert the following:

-25° C = _____ °K

0° C = _____ °K

5° K = _____ °C

290° C = _____ °K

325° K = _____ °C